Practice Redox reactions

$S_2O_3^{2-} + O_2 \rightarrow SO_4^{2-}$ in base

$2\ OH^- + H_2O + S_2O_3^{2-} + 2\ O_2 \rightarrow 2\ SO_4^{2-} + 2\ H_2O$

$VO_2^+ + I_2 \rightarrow VO^{2+} + IO_3^-$ in acid *** CORRECTED FROM THE ORIGINAL***

$8\ H^+ + 10\ VO_2^+ + I_2 \rightarrow 10\ VO^{2+} + 2\ IO_3^- + 4\ H_2O$

ClO_3^- Ý + SO_2Þ $\rightarrow$ ClO_2Þ + SO_4^{2-} Ý in acid

2 ClO_3^- Ý + SO_2Þ $\rightarrow$ 2 ClO_2Þ + SO_4^{2-} Ý

CN^- Ý + OCl^- Ý $\rightarrow$ N_2Þ + HCO_3^- Ý + Cl^- Ý in acid

2 H_2Oà + 4 CN^- Ý + 10 OCl^- Ý $\rightarrow$ 2 N_2Þ + 4 HCO_3^- Ý + 10 Cl^- Ý

MnO_4^- Ý + $Fe(OH)^+$ Ý $\rightarrow$ $Fe(OH)_3$ß + MnO_2ß in base

2 OH^- Ý + 4 H_2Oà + MnO_4^- Ý + 3 $Fe(OH)^+$ Ý $\rightarrow$ 3 $Fe(OH)_3$ß + MnO_2ß + 2H_2Oà