REPORT
MOLECULAR MODELS EXP.

NAME _____

SECTION_____

Compound	Lewis Structure	Central Atom Hybrid	Electron Pair Geometry	Molecular Geometry	Polarity
SCl_2					
PCl_3					
O_3					
SiH_2Cl_2					

REPORT FOR MOLECULAR MODELS EXP. (cont.) NAME _____

Compound	Lewis Structure	Central Atom Hybrid	Electron Pair Geometry	Molecular Geometry	Polarity
PCl_5					
ICl_2^-					
SF_6					
XeF_5^-					

REPORT FOR MOLECULAR MODELS EXP. (cont.) NAME _____

Compound	Lewis Structure	Central Atom Hybrid	Electron Pair Geometry	Molecular Geometry	Polarity
CO					
PCl_3					
SF_4					
NO_3^-					

QUESTIONS FOR MOLECULAR MODELS EXP. NAME _____

1. Answer each of the following for the Nitrate ion (NO_3^-).

   a. Provide the Lewis structures (<u>including</u> resonance forms) for the Nitrate ion. In <u>one</u> of the structures <u>label</u> the hybridization of the Nitrogen and <u>each</u> of the Oxygens.

   b. What are the electron pair <u>and</u> molecular geometries?

      electron pair _____

      molecular _____

   c. Is this a polar or nonpolar ion?

2. The Nitrite ion is slightly different than the Nitrate ion. What are the electron pair and molecular geometries for Nitrite ion? What is its polarity? Include the Lewis structure for this ion in your response.